SRIMAAN COACHING CENTRE-TRICHY-TRB-
POLYTECHNIC LECTURER-PHYSICS-UNIT TEST
QUESTION PAPER-TO CONTACT: 8072230063.

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TRB
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PHYSICS

UNIT TEST QUESTION PAPER-2021-2022
SRIMAAN COACHING CENTRE-TRICHY-TRB-POLYTECHNIC LECTURER-PHYSICS-UNIT TEST QUESTION PAPER-TO CONTACT: 8072230063.

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1. There is no infrared absorption for nitrogen atom because
   a. its polarizability is zero
   b. it has no vibration level
   c. it has no rotation level
   d. its dipole moment is zero

2. In the presence of a magnetic field, a state with total angular momentum quantum number J splits into
   a. J components
   b. 2J components
   c. (2J + 1) components
   d. 2(J + 1) components

3. As the temperature is increased, the intensity of an anti-stokes Raman lines
   a. Increases
   b. Decreases
   c. Unchanged
   d. None of these

4. According to Franck-Condon principle
   a) An electronic transition takes place so rapidly that a rotating molecule does not change its internuclear distance appreciably during transition
   b) The vibration lines in a progression are all of the same intensity
   c) A vibrating molecule does not change its internuclear distance appreciably during an electronic transition
   d) None of these

5. Which one of the following quantities represents \( \frac{hc}{2\pi} \)?
   a. 1973 ÅeV
   b. 973 Å/eV
   c. 938 MeV/Å
   d. 1.973 KeV/Å

6. The hyperfine splitting of atomic levels is due to
   a. Spin - Orbit interaction
   b. Relativistic correction
   c. Electron spin - Nuclear spin interaction
   d. Electron - Electron interaction

7. Electron spin resonance can be exhibited by
   a. Hydrogen atom
   b. Hydrogen molecule
   c. Lithium ion
   d. Nitric oxide (NO) molecule
8. Rayleigh scattering is due to
   a. Fluctuations in absorption
   b. Stimulated emission
   c. Vibrations of $O-H$ bonds
   d. Fluctuations in refractive index

9. The index of refraction of benzene is 1.5. The velocity of light in benzene is,
   a. $1.5 \times 10^8 \text{ m/s}$
   b. $2.0 \times 10^8 \text{ m/s}$
   c. $3.0 \times 10^8 \text{ m/s}$
   d. $4.5 \times 10^6 \text{ m/s}$

10. A spin $j = \frac{3}{2}$ atoms is subjected to a constant external electric fields. The possible schematic energy level is/are
   a. $\frac{3}{2}, \frac{3}{2}, \frac{1}{2}, -\frac{3}{2}$
   b. $\frac{1}{2}, -\frac{1}{2}, \frac{3}{2}, -\frac{3}{2}$
   c. $\frac{3}{2}, \frac{1}{2}, -\frac{1}{2}, -\frac{3}{2}$
   d. None of these

11. The presence of three unpaired electrons is explained by
   a. Hund’s rule
   b. Aufbau principle
   c. Pauli’s exclusion principle
   d. Uncertainty principle

12. The reduced mass ($\mu$) for a system of two particles is
   a. $\frac{m_1m_2}{m_1 + m_2}$
   b. $\frac{m_1m_2}{m_1 - m_2}$
   c. $\frac{1}{m_1} - \frac{1}{m_2}$
   d. Both (a) and (c)

13. A spring of force constant $K$ is cut into three equal parts. Then the force constant of each part is
   a. $K$
   b. $K\frac{1}{3}$
   c. $3K$
   d. $3K^2$

14. The value of proton Bohr magneton is
   a. $5.051 \times 10^{-27} \text{ JT}^{-1}$
   b. $5.051 \times 10^{-27} \text{ Am}^{-1}$
   c. $5.051 \times 10^{-24} \text{ JT}^{-1}$
   d. $5.051 \times 10^{-27} \text{ Am}^2$

15. The spin of $^9\text{F}^{15}$ is
   a. Integral spin
   b. Zero spin
   c. Half integral spin
   d. Zero (or) Integral spin

16. The NMR magnetic field range is..............T
   a. 5 to 10
   b. 2 to 7
   c. 1 to 5
   d. 4 to 8

17. The $^1\text{C}^{12}$ not show any NMR spectrum because
   a. Dipole moment zero
   b. Polarizability zero
   c. Magnetic moment zero
   d. Polarization zero
18. The plasma is cloud of
   a. $\alpha$ - particle           b. Uncharged particles
   c. Neutrino                      d. Completely ionized matter

19. The value of Bohr magneton $m_B$ is
   a. $\frac{e\hbar}{\pi m}$          b. $\frac{e\hbar c}{4\pi m}$
   c. $\frac{e\hbar}{2\pi m}$        d. $\frac{e\hbar}{2m}$

20. The magnetic angle spinning NMR is ...............  
   a. 54.736°    b. 5.4736°
   c. 0.54736°    d. 0°

21. The Bohr magneton is of the order of ...............  
   a. $10^{-23}$ ergs/gauss  b. $10^{-20}$ ergs/gauss
   c. $10^{-25}$ ergs/gauss  d. $10^{-21}$ ergs/gauss

22. Which one of the following molecules does not exhibit a rotational spectra?
   a. $CO$                        b. $HCl$
   c. $H_2$                        d. $H_2O$

23. Infrared absorption can be observed in which of the following molecules?
   a. $N_2$                          b. $O_2$
   c. $C_2$                          d. $HCl$

24. The lowest rotational energy gap in rotational Raman spectrum of linear molecule is ...............  
   a. $2B$                          b. $8B$
   c. $6B$                          d. $4B$

25. All vibrations producing a change in the electric dipole moment of molecule yield ...............  
   a. Raman spectra               b. IR spectra
   c. UV spectra                  d. X-ray spectra

26. The vibrational energy of a molecule in its ground state is
   a. $\frac{3}{2} \hbar \nu_o$       b. $\frac{1}{2} \hbar \nu_o$
   c. $\hbar \nu_o$                 d. 0

27. Raman effect is due to collision of ...............  
   a. Photon with molecule         b. Photon with electron
   c. Electron with atom           d. Electron with proton

28. Near infrared spectrum of diatomic molecule is known as ...............  
   a. Rotational spectrum           b. Vibrational spectrum
   c. Rotational - Vibration spectrum d. Electronic spectrum
29. The total number of electrons in a subshell with orbital quantum number number \( l \) is ................
   a. \( l(l + 1) \)  
   b. \( 2(l + 1) \)  
   c. \( 2(2l + 1) \)  
   d. \( 2l^2 \)

30. The rotational energy levels of a diatomic molecule are ............
   a. continuous  
   b. Discrete and equispaced  
   c. Discrete, but not equispaced  
   d. 2B spaced

31. The rotational energies of a diatomic molecule of rotational constant B joule are ............
   a. B, 2B, 3B,............  
   b. ...0, 2B, 6B, 12B,............  
   c. B, 4B, 9B,............  
   d. 0, 4B, 8B,............

32. The vibration frequency of a diatomic molecule of reduced mass \( (\mu) \) and force constant \( (K) \) ...........
   a. \( -\sqrt{\frac{K}{\mu}} \)  
   b. \( \frac{1}{2\pi} \sqrt{\frac{\mu}{K}} \)  
   c. \( \sqrt{\frac{K}{\mu}} \)  
   d. \( 2\pi \sqrt{\frac{K}{\mu}} \)

33. The rotational spectra of a diatomic molecule falls in the wavelength range ............
   a. 200 Å – 3000 Å  
   b. 1 \( \mu \) – 20 \( \mu \)  
   c. 1 \( \text{Cm} \) – 1 \( \text{m} \)  
   d. 100 \( \mu \) – 500 \( \mu \)

34. In HCl molecule, the energy gap between the two vibrational levels is 0.36 eV. Its zero point energy will be ............
   a. 0  
   b. 0.18 eV  
   c. 0.36 eV  
   d. 0.54 eV

35. Pure vibrational spectrum of a diatomic molecule are when,
   a. It has a centre of symmetry  
   b. It has permanent dipole moment  
   c. It has no magnetic moment  
   d. No electronic transition

36. \( L – S \) coupling occurs often in ............
   a. All atoms  
   b. Lighter atom  
   c. Heavier atom  
   d. Only nuclei

37. Pure rotational spectrum of a diatomic molecule consist of
   a. Two equally spaced lines  
   b. Three equally spaced lines  
   c. Many equally spaced lines  
   d. No regular pattern

38. The separation of the lines in the far infrared spectrum of HBr is 16.90 \( cm^{-1} \), its moment of inertia ............
   a. \( 6.66 \times 10^{-40} \text{ gm cm}^2 \)  
   b. \( 3.30 \times 10^{-38} \text{ gm cm}^2 \)  
   c. \( 3.30 \times 10^{-35} \text{ gm cm}^2 \)  
   d. \( 3.30 \times 10^{-40} \text{ gm cm}^2 \)
39. The typical wavelengths emitted by diatomic molecules in purely vibrational and purely rotational transition are respectively in the region of ...............
a. IR and Visible  
b. Visible and IR  
c. IR and Microwave  
d. Microwave and IR

40. In Born and Oppenheimer, the various form of energies of a molecule are ............... of each other.
a. independent 
b. dependent 
c. equal 
d. Both (a) and (b)

41. Choose the spherical top molecule
a. $BF_3$  
b. $BCl_3$  
c. $CH_3F$  
d. $CH_4$

42. Three moment of inertia are different which molecule represent ...............
a. Linear molecule 
b. Asymmetric molecule 
c. Planar molecule 
d. Symmetric top molecule

43. The solution of the Schrodinger equation for the rigid rotor gives the energy eigen values
a. $\frac{h^2}{2I} \cdot J(J + 1)$  
b. $\frac{h}{8\pi^2 IC} \cdot J(J + 1)$  
c. $B \cdot J(J + 1)$  
d. All of these

44. The zero point energy for anharmonic oscillator
a. $\varepsilon_o = \frac{1}{2} h\nu$  
b. $\varepsilon_o = h\nu$  
c. $\varepsilon_o = \frac{1}{2} \left(1 - \frac{1}{2} \chi_e\right) \nu_e$  
d. $\varepsilon_o = \left(1 - \frac{1}{2} \chi_e\right) \nu_e$

45. The energy of the diatomic molecule obey the ............... 
a. Morse law 
b. Hooke’s law 
c. Schrodinger law 
d. Both (b) and (c)

46. The P.M Morse function is called as ............. function
a. Potential  
b. Vibrational  
c. Rotational  
d. Electrical

47. The Vibrating energy of diatomic molecule are
a. Discrete 
b. Non-Discrete 
c. Quantized 
d. Equal - Unequal

48. The diatomic vibrating rotator of $\Delta J = +1$ which represent ............... branch
a. $O$  
b. $R$  
c. $S$  
d. $P$

49. The fundamental vibration of linear molecule
a. $3N - 6$  
b. $2N - 5$  
c. $3N - 5$  
d. $3N - 1
50. To calculate the fundamental vibration for \( H_2O \) molecule
   a. 4  
   b. 3  
   c. 2  
   d. 6

51. The source of infrared radiation
   a. Block hole  
   b. Sodium lamp  
   c. Nernst glower  
   d. Mercury lamp

52. The Raman scattering are
   a. Coherence  
   b. Incoherence  
   c. Both (a) and (b)  
   d. None of these

53. The Quantum theory of Raman effect the molecule lose energy is
   a. \( h\nu - \Delta E \)  
   b. \( h\nu - \frac{h}{\Delta E} \)  
   c. \( h\nu + \Delta E \)  
   d. \( h\nu + \frac{h}{\Delta E} \)

54. The selection rule for rotational Raman spectrum is
   a. \( \pm 1 \)  
   b. \( \pm 1, \pm 2, \pm 3, \ldots \)  
   c. \( 0, \pm 1 \)  
   d. \( 0, \pm 2 \)

55. The homodiatomic molecule gives ............. spectrum
   a. Rotational  
   b. Pure vibrational  
   c. Pure rotational Raman  
   d. None of these

56. The \( CO_2 \) molecule of asymmetric stretching mode of spectrum of Raman and IR is
   a. Active and Inactive  
   b. Inactive and Active  
   c. Active and Active  
   d. Inactive and Inactive

57. In electronic spectra the rotational changes produce rotational ............. structure
   a. fine  
   b. coarse  
   c. hyperfine  
   d. splitting

58. In the Fronk-Condon principle intensity of vibrational electronic spectra energy curve is
   a. Hyperbolic  
   b. Periodic  
   c. Parabolic  
   d. Non-parabolic

59. In Fronk - Condon principle parabola are referred to as ............. parabola
   a. Hyperbolic  
   b. Fortrat  
   c. Periodic  
   d. Elliptical

60. The predissociation is explained by
   a. Hyperbolic curve  
   b. Morse curve  
   c. Circle  
   d. None of these
61. In a magnetic field splits up into ........... state of NMR
   a. $2I + 1$  
   b. $2(2I + 1)$  
   c. $J + 1$  
   d. $2(J + 2)$

62. In NMR the powerful tool for structure determination are
   a. $B, \delta$  
   b. $\tau, \nu$  
   c. $\sigma, \tau$  
   d. $\tau, \delta$

63. When placed in a magnetic field the surrounding electron cloud tends to circulate in such a direction as to produce a field opposing that applied is called
   a. Diamagnetic  
   b. Paramagnetic  
   c. Ferromagnetic  
   d. Ferrimagnetic

64. Calculate the strength of the magnetic field required to give a precessional frequency of 100 MHz for $^{19}$F nucleus $g_N = -0.757$, $\mu_N = 5.051 \times 10^{-27}$ JT$^{-1}$, $I = 5/2$. The separation between any to adjacent is $g_N\mu_NB_o$.
   a. 17.13 T  
   b. 17.33 T  
   c. 17.63 T  
   d. 17.93 T

65. What is the nuclear $g_N$ factor for $^{19}$F nucleus which has a magnetic moment of 2.6273 $\mu_N$, $I = 1/2$.
   a. 4.2546  
   b. 25.321  
   c. 55.231  
   d. 5.25

66. The ESR arising from
   a. Parallel electron  
   b. Unpaired electron  
   c. Same electron  
   d. Equal mass of electron

67. The resonance condition for ESR
   a. $E = g\mu_B B$  
   b. $E = g\mu_N B$  
   c. $E = g\mu_N T$  
   d. $E = \mu_B B$

68. The electron-electron coupling is called
   a. Hyperfine structure  
   b. Fine structure  
   c. Resonance structure  
   d. Nuclear structure

69. A free electron is placed in a magnetic field of strength 1.3 T. Calculate the resonance frequency if $g = 2.0023$.
   a. 3.643 GHz  
   b. 364.3 GHz  
   c. 36.43 GHz  
   d. 30 GHz

70. The quadrupole nucleus having spin
   a. $I = 0$  
   b. $I = 1/2$  
   c. $I \geq \frac{1}{2} + 1$  
   d. $I > \frac{1}{2}$
71. $I \geq 1$ and $eQ > 0$ which represent the nucleus

72. NQR is observed for
   a. Gas  b. Liquid  c. Solid  d. Plasma

73. The Mössbauer spectrum is called as
   a. $\alpha$-ray  b. $\beta$-ray  c. $\gamma$-ray  d. $X$-ray

74. The exciting line in an experiment is $5460 \text{ Å}$ and the stokes line is at $5520 \text{ Å}$. Find the wave number displacement
   a. $199 \text{ cm}^{-1}$  b. $195 \text{ cm}^{-1}$  c. $205 \text{ cm}^{-1}$  d. $299 \text{ cm}^{-1}$

75. In NMR spectrum chemical shift measuring unit is
   a. $npm$  b. $mpn$  c. $ppm$  d. $pµm$

CONTINUATION OF THE QUESTIONS.....

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SRIMAAN COACHING CENTRE-TRICHY.

TO CONTACT: 8072230063.
# SRIMAAN COACHING CENTRE-TRICHY.

**TO CONTACT:** 8072230063.

## ANSWER KEY

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**TO CONTACT:** 8072230063.

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SRIMAAN COACHING CENTRE
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